

El machismo camaleón: The Nobel Committee Case

VMOG • 07/03/2019



Visión general

Premio Nobel

Desde 1901 en 5 categorías

Espanoles

8: 6 literatura y 2 medicina

Mujeres

51/853=6%



Medicina 1959: Síntesis biológica del ARN y del ADN

11 November 1955, Volume 122, Number 3176

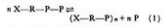
SCIENCE

Enzymatic Synthesis of Nucleic Acidlike Polynucleotides

Marianne Grunberg-Manago, Priscilla J. Ortiz, Severo Ochoa

The mechanisms of synthesis of the polynucleotide chains of nucleic acids have remained obscure despite notable advances in our understanding of the enzymatic mechanisms involved in the synthesis of the mononucleotides, the purine and pyrimidine bases, and the sugar moieties.

As briefly reported in a recent note from this laboratory (1), an enzyme isolated from the microorganism *Azotobacter vinelandii* catalyzes the synthesis of polynucleotides from 5'-nucleotide diphosphates with release of orthophosphate. The reaction requires magnesium ions and is reversible. The available evidence indicates that the *Azotobacter* enzyme catalyzes the reaction



where R stands for ribose, P—P for pyrophosphate, P for orthophosphate, and X for one or more of the following bases: adenine, hypoxanthine, guanine, uracil, or cytosine.

Chemical and enzymatic degradation of the biosynthetic polynucleotides showed that they are made up of 5'-mononucleotide units linked to one another through 3'-phosphoribose ester bonds as in RNA (2). Thus, in analogy with polychaetins, reversible phosphorylation may be a major mechanism in the biological breakdown and synthesis of polynucleotide chains. For this reason, the name polynucleotide phosphorylase has been proposed (1) for the new en-

zyme. Evidence that this enzyme brings about the synthesis of RNA-like polynucleotides is presented in this article (3).

Polynucleotide Phosphorylase

The enzyme was discovered in the course of a study of biological phosphorylation mechanisms (4) when it was found that, in the presence of Mg^{2+} , *Azotobacter* extracts catalyzed an exchange of ^{32}P -labeled orthophosphate with the terminal phosphate groups of the 5'-nucleotide diphosphates of adenosine, inosine, guanosine, uridine, and cytidine. Following partial purification of the activity, there was no reaction with 5'-nucleotide mono- or triphosphates such as AMP, ATP, IMP, or ITP (2). The "exchange" was accompanied by the liberation of orthophosphate, as one would expect from the reaction of Eq. 1. With mixtures of 5'-nucleotide diphosphates, radioactive phosphate was incorporated in all of them.

As previously reported (1), the enzyme activity was purified about 40-fold through ammonium sulfate fractionation and calcium phosphate gel adsorption steps; the rate of the ADP-orthophosphate exchange was employed as an assay. The ratio of the rates of the ADP-orthophosphate exchange to orthophosphate liberation remained constant on purification, suggesting that the two activities were related to each other.

Single Polymer

On incubation of purified polynucleotide phosphorylase with nucleotide di-

phosphates in the presence of Mg^{2+} , there is a disappearance of nucleotide diphosphate with liberation of a stoichiometric amount of orthophosphate. The reaction reaches equilibrium about 50 to 60 per cent of the nucleotide diphosphate has disappeared. The disappearing diphosphate is converted into a polynucleotide, as is borne out by the following facts. (1) The newly formed compound is strongly negatively charged, for it is retained by Dowex-1 anion exchange columns (5) following elution with formic acid at concentrations higher than those required to elute the most acidic mononucleotides. (2) The product is nondialyzable against distilled water or dilute salt solutions and is quantitatively precipitated by trichloroacetic acid in alcohol in the cold. It can be isolated in this way. (3) It is soluble in water, giving more or less viscous solutions with a typical nucleotide ultraviolet absorption spectrum. (4) It remains at the origin of paper chromatograms, whichever the solvent system used, and cannot be eluted with water. Single polymers containing AMP, IMP, GMP, UMP, or CMP as the only basic unit have been obtained by incubating polynucleotide phosphorylase with the corresponding 5'-nucleotide diphosphates (6). If the use of heat and/or acid is avoided in the isolation of the polymers, their average molecular weights may be very high. Values of 570,000 and 800,000, respectively, were obtained by light scattering for the AMP and IMP polymers (7).

Mild alkaline hydrolysis (8) of the IMP polymer—that is, the polynucleotide formed from IDP—yielded an approximately equimolar mixture of 2'-IMP and 3'-IMP. These products were identified by paper chromatography with the solvent system No. 3 of Markham and Smith (9). As is well known, RNA yields mixtures of the 2'- and 3'-mononucleotides on mild hydrolysis with alkali (10). The chromatographic identification of the alkaline hydrolysis products of the IMP polymer as 2'- and 3'-IMP has been confirmed by the following experiments (Table 1). (1) About 80 per cent of the product of alkaline hydrolysis of the IMP polynucleotide is hydrolyzed by 1.0N hydrochloric acid in 20 minutes. It will be observed that, under these conditions, both 2'-IMP and 3'-IMP are hydrolyzed to a similar extent—that is, by

Enzymatic Synthesis of Nucleic Acidlike Polynucleotides.
Science 122 (1955) 907.

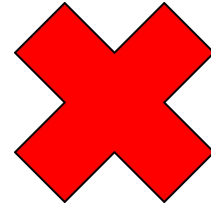
M. Grundberg-Manago, P. J. Ortiz y S. Ochoa.

Dr. Ochoa is chairman and Miss Ortiz is research assistant of the Department of Biochemistry, The New York University College of Medicine. Dr. Grundberg-Manago held a fellowship in the department while she was on leave of absence from the Institut de Biologie Pasteur-Clayton, Paris.

Medicina 1959: Síntesis biológica del ARN y del ADN



Severo Ochoa



Marianne Grundberg-Manago

Medicina 1959: Síntesis biológica del ARN y del ADN

— — —

Curiosidad/maldad: El resultado por el que le dieron el Nobel a Severo Ochoa se demostró ser **falso** pocos años después: la enzima descubierta, la **polinucleótido fosforilasa**, no sintetizaba ARN, proceso realizado por la ARN polimerasa. Así lo reconoció la academia sueca años después. No obstante, su descubrimiento sirvió para realizar otro más importante aún: el descifrado del código genético. Este descubrimiento volvió a valer un premio Nobel, pero en este caso el **olvidado** fue Severo Ochoa: ¿Cuestión de karma?

Física 1957: Leyes de paridad en partículas elementales

The branching ratio of the two modes of decay of Fm^{253} , i.e., E.C./ α , was found to be about 8.5—which gives $\sim 89.5\%$ decay by electron capture and $\sim 10.5\%$ by alpha emission. It was not possible to measure the cross section for the $\text{Cf}^{252}(\alpha, n)\text{Fm}^{253}$ reaction because Fm^{253} could also be produced from other californium isotopes in the target.

A previous publication¹ on a possible identification of the Fm^{253} gave the values of 6.85 ± 0.04 Mev for the alpha-particle energy, and a half-life > 10 days.

It is a pleasure to thank the crew of the 60-inch cyclotron for their extremely careful and skillful operation of the machine during the bombardment. We wish to thank Professor Glenn T. Seaborg for his continued interest.

¹ On leave from the Israel Atomic Energy Commission, Weizmann Institute of Science, Rehovoth, Israel.

² Thompson, Gléason, Harvey, and Choppin, *Phys. Rev.* **83**, 908 (1954).

³ Harvey, Chatham-Strode, Gléason, Choppin, and Thompson, *Phys. Rev.* **104**, 1315 (1958).

⁴ Thompson, Harvey, Choppin, and Seaborg, *J. Am. Chem. Soc.* **76**, 4229 (1954); Choppin, Harvey, and Thompson, *J. Inorg. and Nuclear Chem.* **2**, 60 (1956).

⁵ Friedland, Gandler, Baccus, Spöck, and Fildis, *Phys. Rev.* **102**, 353 (1956).

Experimental Test of Parity Conservation in Beta Decay*

C. S. Wu, *Columbia University, New York, New York*

AND

Y. S. Wu, H. W. Haggard, T. D. Hoopes, and R. P. Hulsizer,

National Bureau of Standards, Washington, D. C.

(Received January 15, 1957)

IN a recent paper¹ on the question of parity in weak interactions, Lee and Yang critically surveyed the experimental information concerning this question and reached the conclusion that there is no existing evidence either to support or to refute parity conservation in weak interactions. They proposed a number of experiments on beta decays and hyperon and meson decays which would provide the necessary evidence for parity conservation or nonconservation. In beta decay, one could measure the angular distribution of the electrons coming from beta decays of polarized nuclei. If an asymmetry in the distribution between θ and $180^\circ - \theta$ (where θ is the angle between the orientation of the parent nuclei and the momentum of the electrons) is observed, it provides unequivocal proof that parity is not conserved in beta decay. This asymmetry effect has been observed in the case of oriented Co^{60} .

It has been known for some time that Co^{60} nuclei can be polarized by the Rose-Gorter method in cerium magnesium (rock) nitrate, and the degree of polarization detected by measuring the anisotropy of the succeeding gamma rays.² To apply this technique to the present problem, two major difficulties had to be over-

come. The beta-particle counter should be placed inside the demagnetization cryostat, and the radioactive nuclei must be located in a thin surface layer and polarized. The schematic diagram of the cryostat is shown in Fig. 1.

To detect beta particles, a thin anthracene crystal $\frac{1}{2}$ in. in diameter $\times \frac{1}{4}$ in. thick is located inside the vacuum chamber about 2 cm above the Co^{60} source. The scintillations are transmitted through a glass window and a Lucite light pipe 4 feet long to a photomultiplier (6292) which is located at the top of the cryostat. The Lucite lens is machined in a logarithmic spiral shape for maximum light collection. Under this condition, the Co^{60} conversion line (624 kev) still retains a resolution of 17%. The stability of the beta counter was carefully checked for any magnetic temperature effects and none were found. To measure the amount of polarization of Co^{60} , two additional NaI gamma scintillation counters were installed, one in the equatorial plane and one near the polar position. The observed gamma-ray anisotropy was used as a measure of polarization and, effectively, temperature. The tank susceptibility was also monitored but this is of secondary significance due to secondary magnetic effects, and the gamma-ray anisotropy alone provides a reliable measure of nuclear polarization. Specimens were made by taking good single crystals of cerium magnesium nitrate and growing on the upper surface only an additional crystalline layer containing Co^{60} . One might point out here that since the beta decay of Co^{60} involves a change of spin of

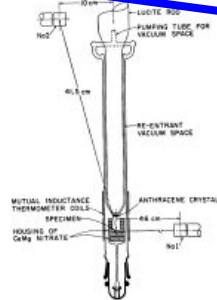


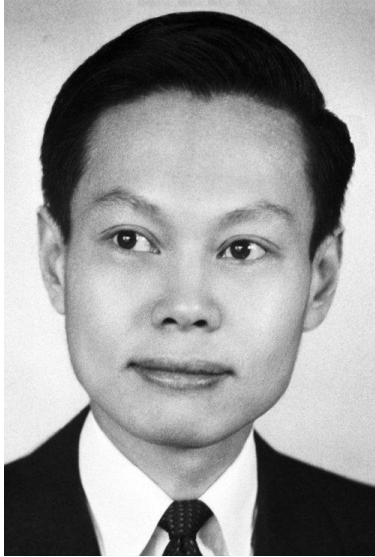
FIG. 1. Schematic drawing of the lower part of the cryostat.

Experimental test on parity conservation in beta decay.

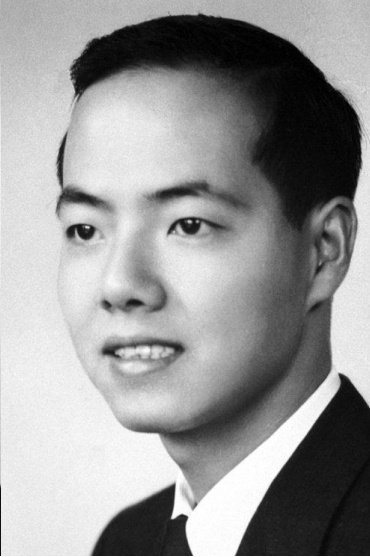
Phys. Rev. Lett. 105 (1957) 1413.

C. S. Wu.

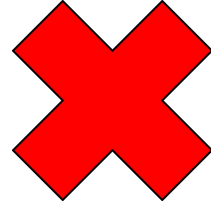
Física 1957: Leyes de paridad en partículas elementales



Chen Ning
Yang



Tsung Dao
Lee



Chien Siung
Wu

Medicina 192: Estructura molecular del ADN

King's College, London. One of us (J. D. W.) has been aided by a fellowship from the National Foundation for Infectious Diseases.

J. D. WATSON
F. H. C. CRICK

Medical Research Council Unit
Study of the Molecular Structure of
Biological Systems,
Cavendish Laboratory, Cambridge,
April 2.

¹Pauling, L., and Corey, R. B., *Kevex* 172, 143 (1952); *Proc. U.S. Nat. Acad. Sci.*, 38, 55 (1952).
²Pauling, L., *J. Am. Chem. Soc.*, 74, 354 (1952).
³Chargaff, E., *in* *Progress in Nucleic Acid Research*, Vol. 1, and *Ann. N.Y. Acad. Sci.*, 56, 476 (1951).
⁴Arnold, C., *Ann. Soc. Exp. Biol.*, 1, *Statist. Anal.*, 56 (1948).
⁵Wilkins, M. H. F., and Randall, J. T., *Nucleic Acids Respt.*, 4, 318 (1952).

MOLECULAR STRUCTURE OF NUCLEIC ACIDS

A Structure for Desoxyribose Nucleic Acid

WHAT IS the structure of the salt of desoxyribose nucleic acid (D.N.A.)? This structure has novel features which are of considerable biological interest.

A structure for nucleic acid has already been proposed by Pauling and Corey¹. They kindly made their manuscript available to us in advance of publication. Their model consists of three inter-twined chains, with the phosphates near the fibre axis, and the bases on the outside. In our opinion, this structure is unsatisfactory for two reasons: (1) We believe that the material which gives the X-ray diagrams is the salt, not the free acid. Without the acidic hydrogen atoms it is not clear what forces would hold the structure together, especially as the negatively charged phosphates near the axis will repel each other. (2) Some of the van der Waals distances appear to be too small.

Another three-chain structure has also been suggested by Fraser (in press). In his model the phosphates are on the outside and the bases on the inside, linked together by hydrogen bonds. This structure as described is rather ill-defined, and for this reason we shall not comment on it.

We wish to put forward a radically different structure for the salt of desoxyribose nucleic acid. This structure has two helical chains each coiled round the same axis (see diagram). We have made the usual chemical assumptions, namely, that each chain consists of phosphate di-ester groups joining 3'-O-desoxy-ribose nucleoside units with 2',5' linkages. The two chains differ in that their bases are related by a dyad perpendicular to the fibre axis. Both chains follow right-handed helices, but owing to the dyad the sequence of the atoms in the two chains run in opposite directions. Each chain loosely resembles Furberg's model No. 1; that is, the bases are on the inside of the helix and the phosphates on the outside. The configuration of the sugar and the atoms near it is close to Furberg's 'standard configuration', the sugar being roughly perpendicular to the attached base. There

is a residue on each chain every 3.4 Å, in the 2-dimension. We have assumed an angle of 36° between adjacent residues in the same chain, so that the structure repeats after 10 residues on each chain, that is, after 34 Å. The distance of a phosphate atom from the fibre axis is 10 Å. As the phosphates are on the outside, reactions have easy access to them.

The structure is an open one, and its water content is rather high. At lower water contents we would expect the bases to fall so that the structure could become more compact.

The novel feature of the structure is the manner in which the two chains are held together by the purine and pyrimidine bases. The phosphate groups are perpendicular to a single base from one chain being hydrogen-bonded to a single base from the other chain, so that the two helix rise by sides with identical *h*-*h*-*h* orientation. One of the pair must be a purine and the other a pyrimidine for bonding to occur. The hydrogen bonds are made as follows: purine position 1 to pyrimidine position 1; purine position 6 to pyrimidine position 6.

If it is assumed that the bases only occur in the structure in the most plausible tautomeric form (that is, with the keto rather than the enol configuration) it is found that only specific pairs of bases can form together. These pairs are: adenine (purine) with thymine (pyrimidine), and guanine (purine) with cytosine (pyrimidine).

In other words, if an adenine forms one member of a pair, on either chain, then on these assumptions the other member must be thymine; similarly for guanine and cytosine. The sequence of bases on a single chain does not appear to be restricted in any way. However, if only specific pairs of bases can be formed, it follows that if the sequence of bases on one chain is given, then the sequence on the other chain is automatically determined.

It has been found experimentally^{2,3} that the ratio of the amounts of adenine to thymine, and the ratio of guanine to cytosine, are always very close to unity for desoxyribose nucleic acid.

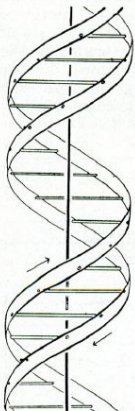
It is probably impossible to build this structure with a ribose sugar in place of the desoxyribose, as the extra oxygen atom would make too close a van der Waals contact.

The previously published X-ray data^{4,5} on desoxyribose nucleic acid are insufficient for a rigorous test of our structure. So far as we can tell, it is roughly compatible with the experimental data, but it must be regarded as unproved until it has been checked against more exact results. Some of these are given in the following communications. We were not aware of the details of the results presented there when we devised our structure, which rests mainly though not entirely on published experimental data and stereochemical arguments.

It has not escaped our notice that the specific pairing we have postulated immediately suggests a possible copying mechanism for the genetic material.

For details of the structure, including the conditions assumed in building it, together with a set of co-ordinates for the atoms, will be published elsewhere.

We are much indebted to Dr. Jerry Donohue for constant advice and criticism, especially on interatomic distances. We have also been stimulated by a knowledge of the general nature of the unpublished experimental results and those of Dr. M. H. F. Wilkins, Dr. R. E. Franklin and their co-workers at



This figure is purely diagrammatic. The two ribbons symbolize the two phosphate-sugar chains, and the horizontal rods the pairs of bases holding the chains together. The vertical line marks the fibre axis.

A structure for desoxyribose nucleic acid
Nature 171 (1953) 737.

J. D. Watson, F. H. C. Crick.

This figure is purely diagrammatic.
They forgot to add "and it is done by my wife".

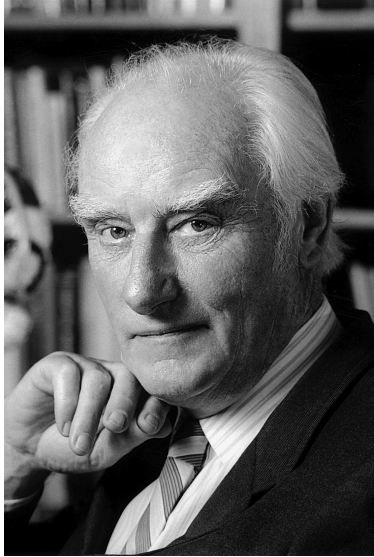
Medicina 1962: Estructura molecular del ADN

* The information reported in this section was very kindly reported to us prior to its publication by Drs Wilkins and Franklin. We are most heavily indebted in this respect to the King's College Group, and we wish to point out that without this data the formulation of our structure would have been most unlikely, if not impossible. We should at the same time mention that the *details* of their X-ray photographs were not known to us, and that the formulation of the structure was largely the result of extensive model building in which the main effort was to find any structure which was stereochemically feasible.

The complementary structure
of deoxyribonucleic acid.
Proc R Soc A 223 (1954) 80

Curiosidad: Durante el desarrollo de su modelo, Watson y Crick nunca tocaron o miraron directamente una fibra de ADN. Su artículo no tiene ni un experimento, y consiste en especulaciones basadas en experimentos de otros científicos. En un pie de página de un artículo en 1954 Crick y Watson escribieron, “we wish to point out that without those data the formulation of our structure would have been most unlikely, if not impossible”. En 1999 Watson admitió, “the Franklin photograph was the key event”.

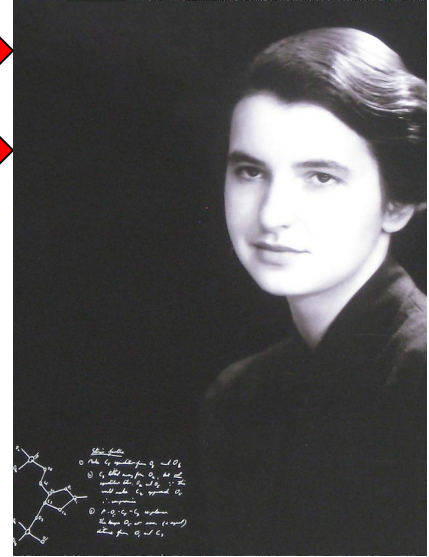
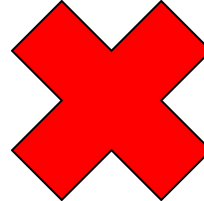
Medicina 1962: Estructura molecular del ADN



Francis Harry
Crick



James Dewey
Watson



Rosalind Franklin

Física 1974: Descubrimiento de los púlsares

Ph.D. Dissertation
6567

THE MEASUREMENT OF RADIO SOURCE DIAMETERS
USING A DIFFRACTION METHOD

The measurement of radio source diameters
using a diffraction method.
PhD. Dissertation. (1964)
<https://doi.org/10.17863/CAM.4926>

THE BOARD OF GRADUATE STUDIES
APPROVED THIS DISSERTATION
FOR THE Ph. D. DEGREE ON 18 FEB 1969

S. J. Bell

SUSAN JOCKELYN BELL

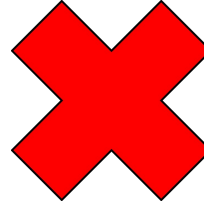
New Hall

September 1968

Física 1974: Descubrimiento de los púlsares



Antony
Hewish



Susan Jocelyn
Bell

Conclusiones

1. Sexo y orientación sexual **no** son **discutibles**
2. El **género** es un constructo social
3. El trabajo de la mujer en ciencia ha sido **minorizado** porque la mujer está minorizada
4. Toda información ha de estar basada en datos que se puedan **contrastar**
5. Soy un **profesor** guay

Fin