

CALCULA O NÚMERO DE OXIDACIÓN DE ÁTOMO



$$n^{\circ} \text{oxidi} = 0$$



$$n^{\circ} \text{ox} = 0$$

MAZCARO



$$n. \text{ox} = 0$$

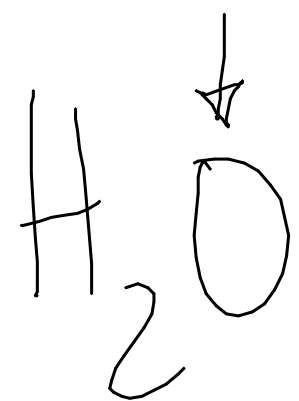


$$n^{\circ} \text{ox} = 0$$

número oxidación



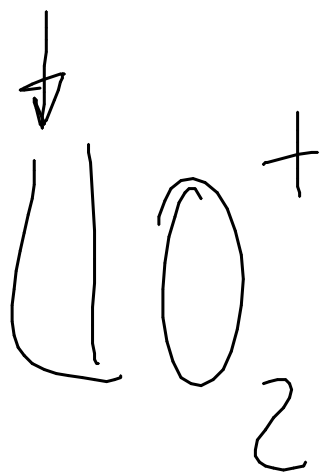
$$n^{\circ} \text{ox} = 0$$



$$\frac{q=0}{\text{---}}$$

$$2(+1) + X = 0$$

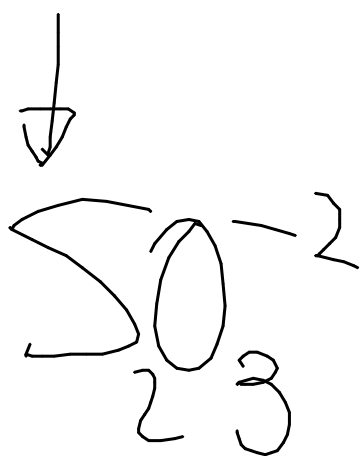
$$X = -2 \rightarrow \text{n}^\circ \text{ox do O} = -2$$



$$\boxed{q=1}$$

$$X + 2(-2) = 1$$

$$\text{n}^\circ \text{ox Cl} = +5$$

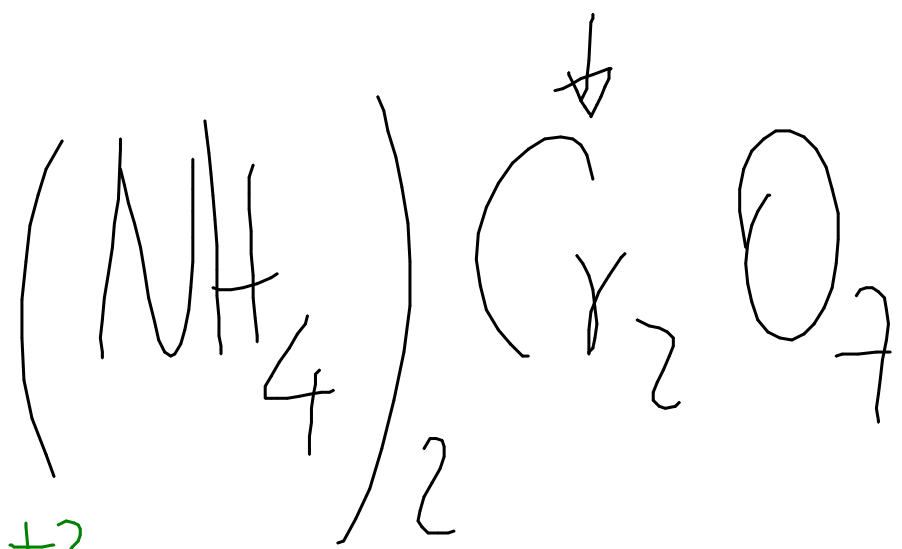


$$\boxed{q=-2}$$

$$2X + 3(-2) = -2$$

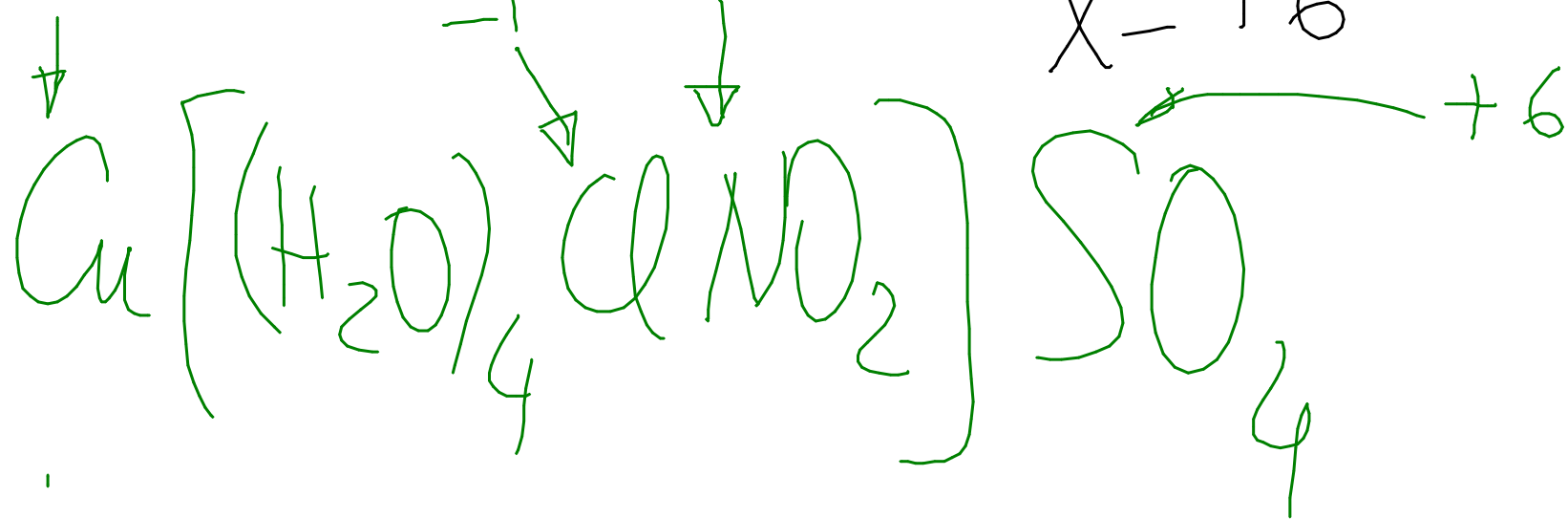
$$X = +2 \text{ n}^\circ \text{ox do S} = +2$$

$$\boxed{9=0}$$



$$2(-3) + 8(+1) + 2X + 7(-2) = 0$$

+2

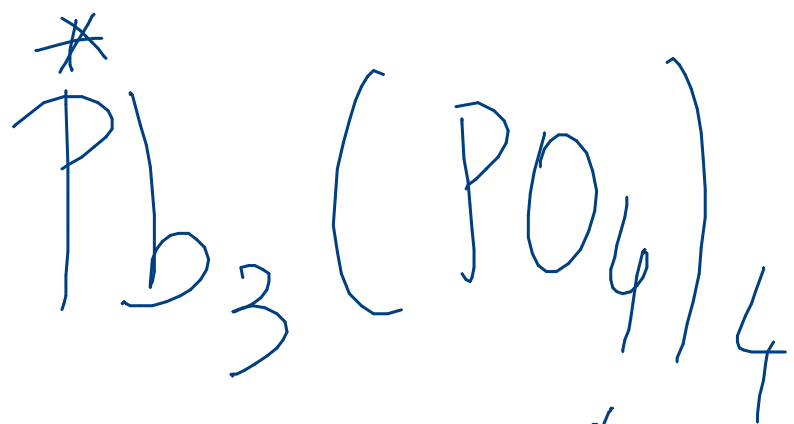


$$X = +6$$

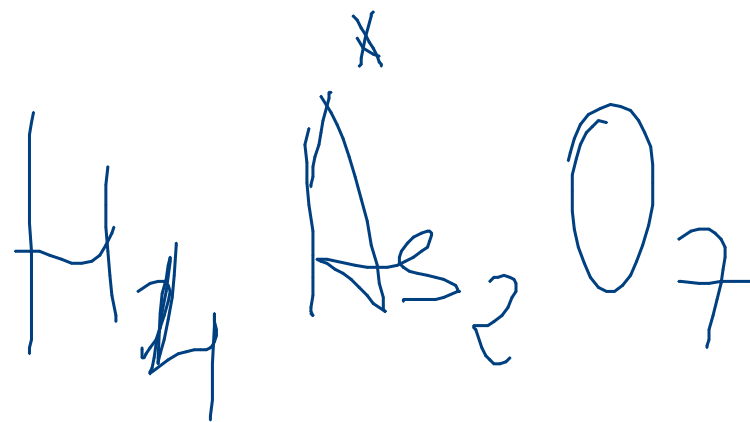
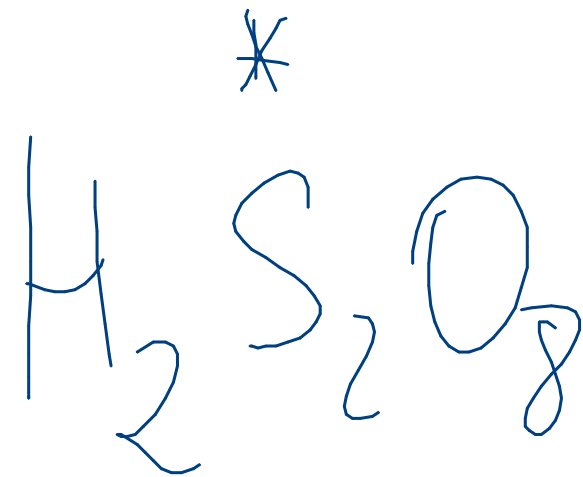
+6

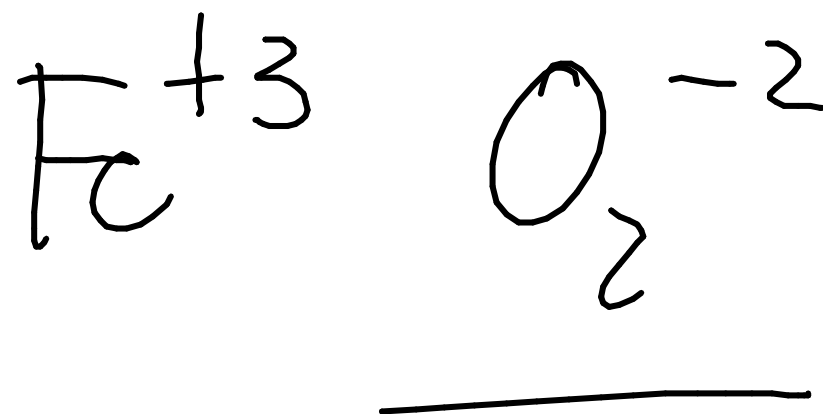
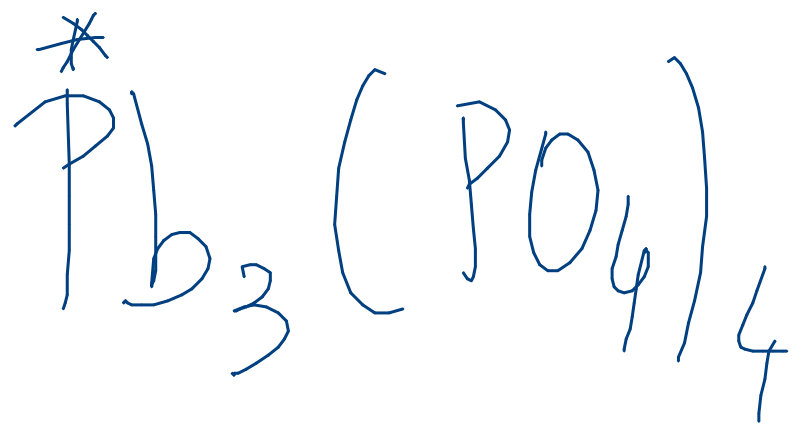
$$+2 + 8(+1) + 10(-2) + (-1) + X + 6 = 0$$

$$X = +5$$

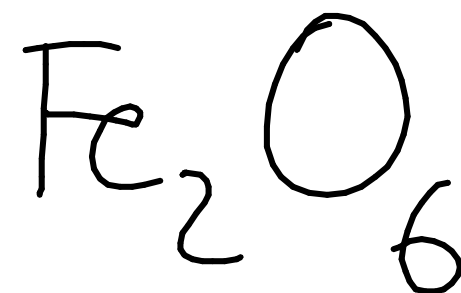


$$\text{OSP} = +5^-$$

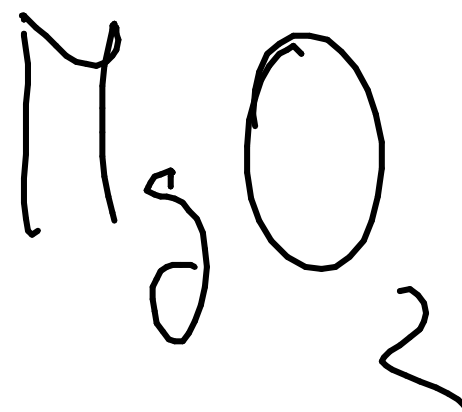




$$3x + 4(5) + 16(-2) = 0$$



$$3x + 20 - 32 = 0$$

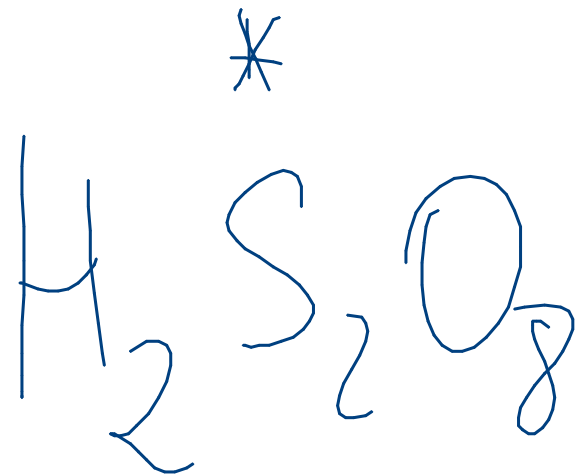


$$x = 4$$



$$3(-3) + 12(1) + (+1) + 2X + 6(-2) = 0$$

$$X = +4$$



$$2(1) + 2x + 8(-2) = 0$$

$$2x = 14$$

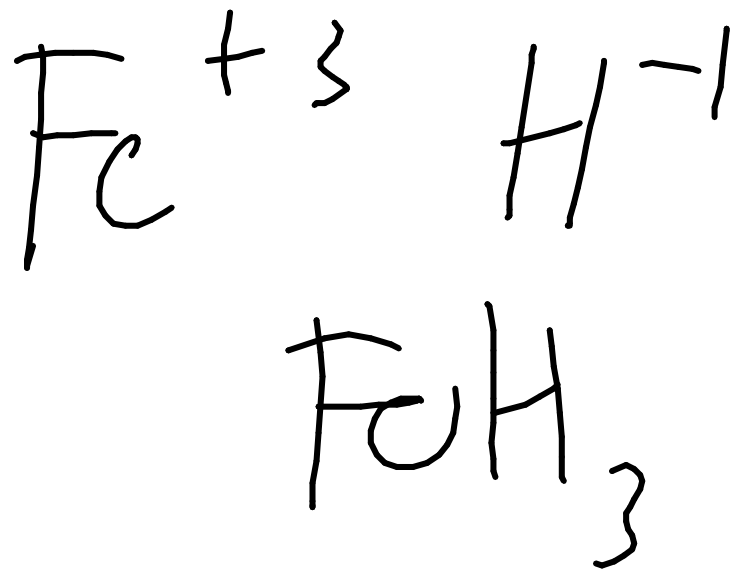
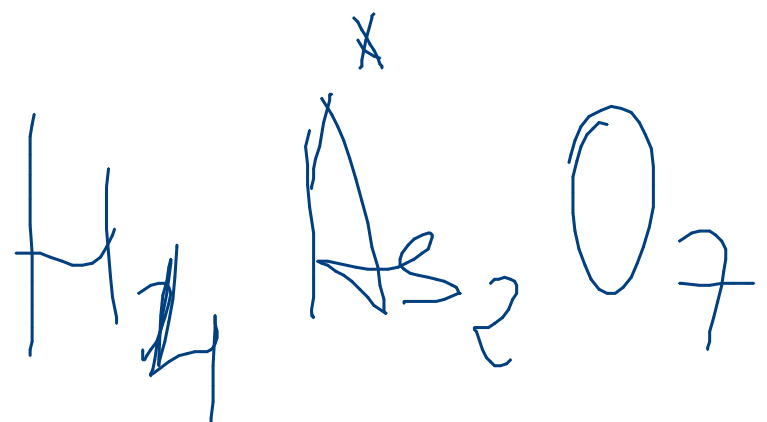
$$x = 7$$

$$* \quad \text{Al}(\text{FeO})_{2/3}$$

$$2x + 3(+2) + \cancel{6}(-2) = 0$$

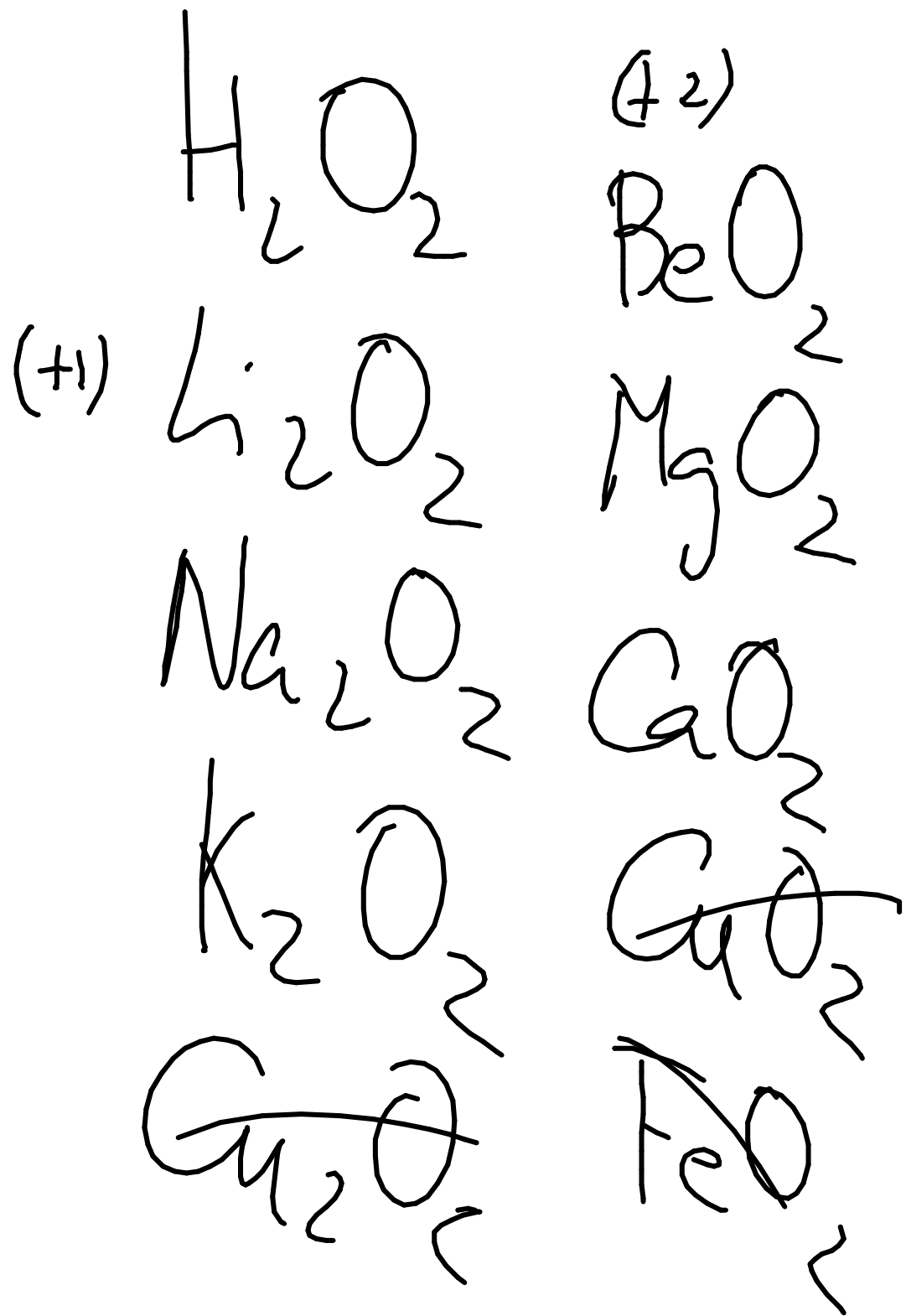
$$2x = 6$$

$$x = +3$$



$$4(1) + 2x + 7(-2) = 0$$

$$x = +5$$



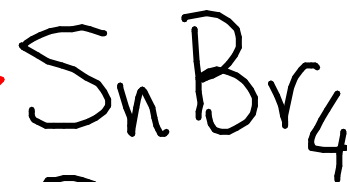
Peróxidos más
comúns

1 Formula:

a) Fluoruro de bario



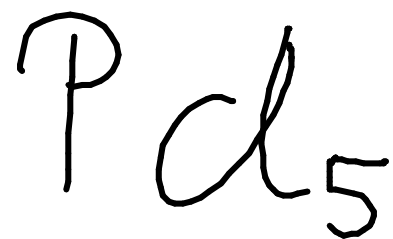
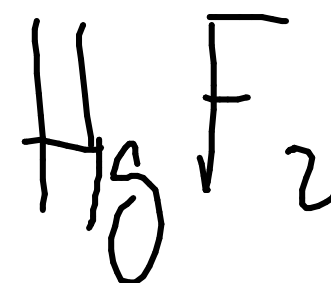
b) Tetrabromuro de estaño



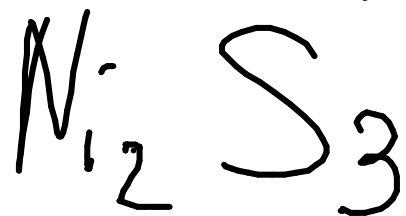
c) Trifluoruro de fósforo



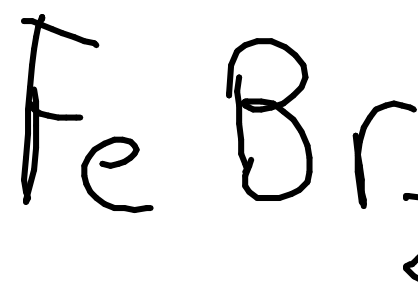
d) Difluoruro de mercurio



e) Pentacloruro de fósforo



f) Sulfuro de níquel(III)



g) Bromuro de hierro(II)



h) Nitruro de boro

2 Nombra:

- a) PbF_2
- b) CsBr
- c) AlCl_3
- d) CaI_2

- e) Cd_2S_3
- f) Li_3N
- g) SrSe
- h) Na_2S

- a) **Fluoruro de Plomo (II)**
- b) **Bromuro de Cesio**
- c) **Cloruro de Aluminio**
- d) **Ioduro de Calcio**
- e) **Sulfuro de Cadmio (III)**
- f) **Nitruro de Litio**
- g) **Seleniuro de Estroncio**
- h) **Sulfuro de Sodio**

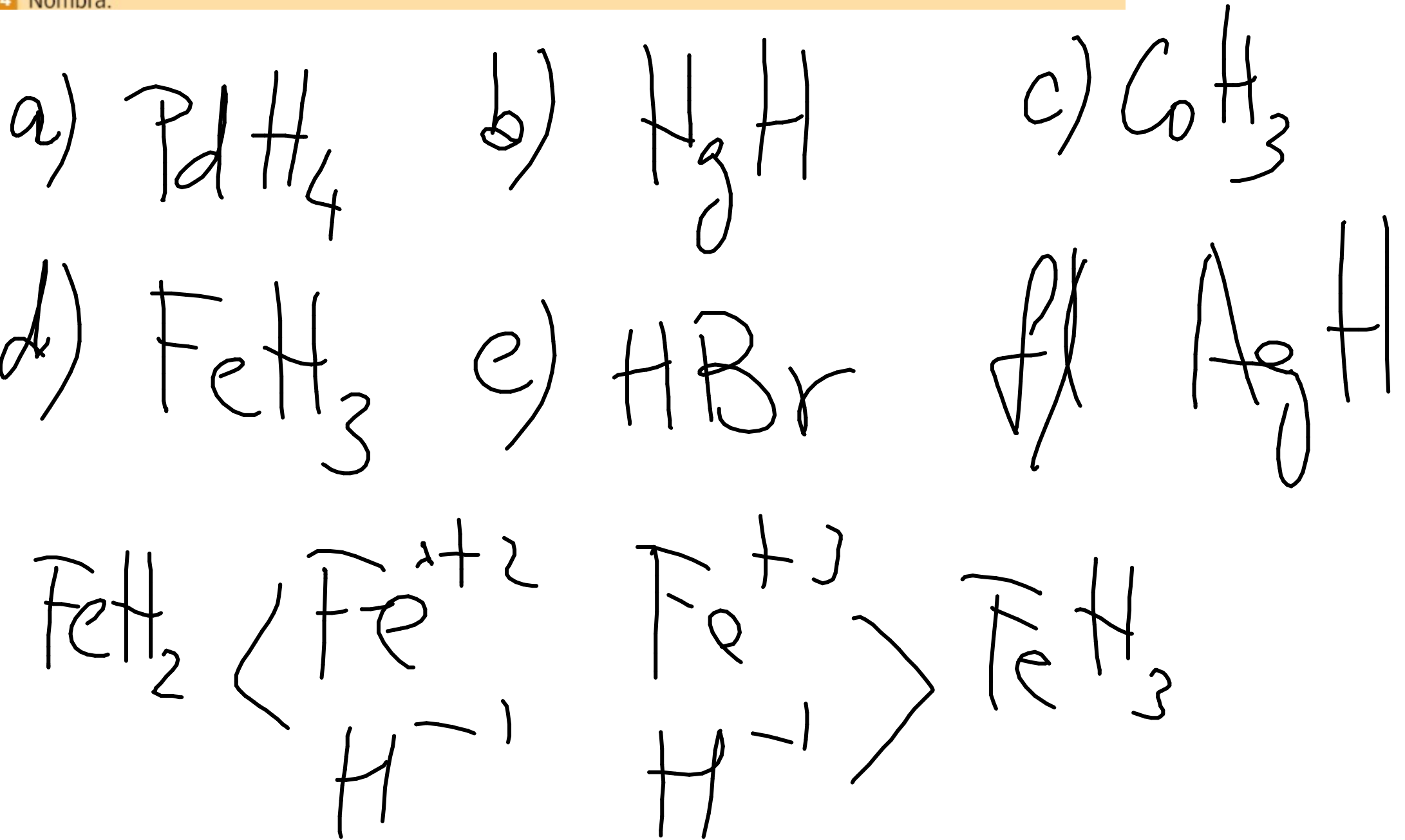
3 Formula:

- a) Tetrahidruro de paladio
- b) Monohidruro de mercurio

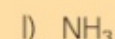
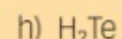
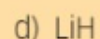
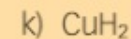
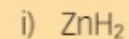
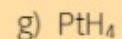
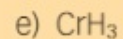
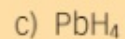
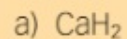
- c) Hidruro de cobalto(III)
- d) Trihidruro de hierro

- e) Bromuro de hidrógeno
- f) Hidruro de plata

4 Nombra:



4 Nombra:



5 Formula:

a) Hidruro de Calcio

b) acido iodhídrico ou ioduro de hidróxeno

c) Hidruro de plomo (IV) ou tetrahidruro de chumbo (plomo)

d) Hidruro de litio

e) trihidruro de cromo ou hidruro de cromo (III)

f) ac. sulfhídrico ou sulfuro de hidróxeno

g) tetrahidruro de platino ou hidruro de platino

h) ac. telurhídrico ou telururo de hidróxeno

i) hidruro de zinc

j) ac. bromhídrico ou bromuro de hidróxeno

k) hidruro de cobre (II) ou dihidruro de cobre

l) amoníaco

5 Formula:

a) Metano

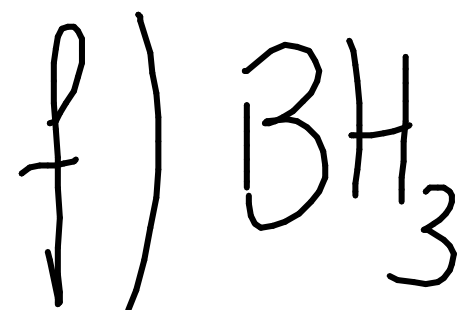
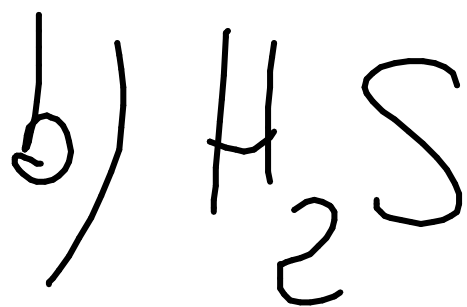
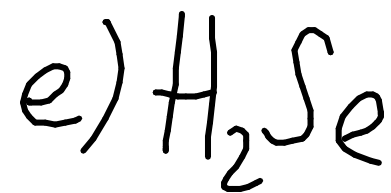
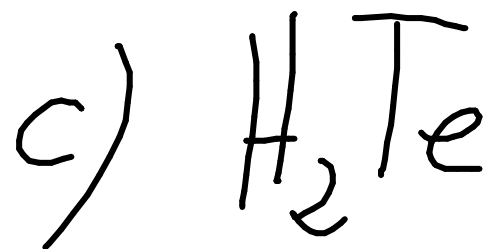
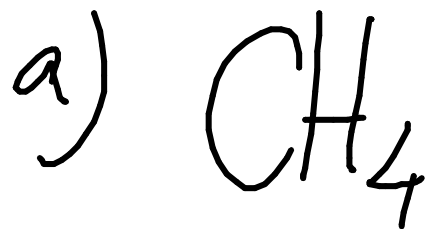
b) Ácido sulfhídrico

c) Ácido telurhídrico

d) Fosfano

e) Ácido selenhídrico

f) Borano



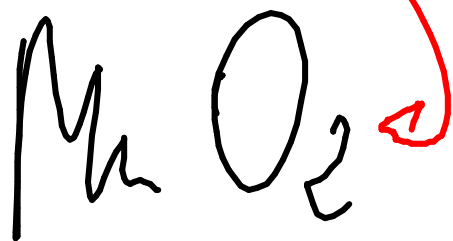
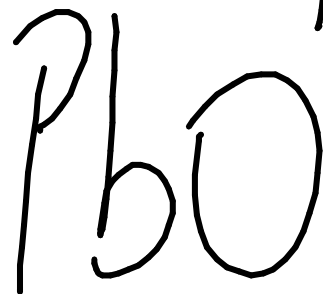
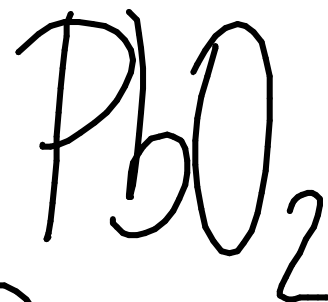
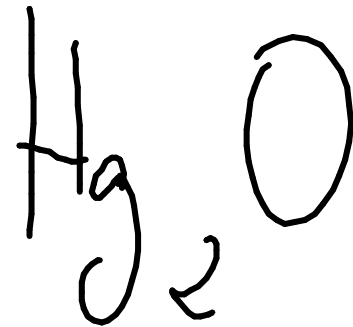
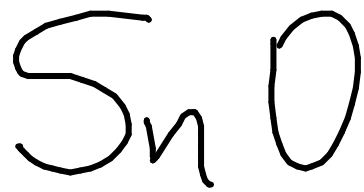
6 Nombra con el nombre sistemático y el tradicional:

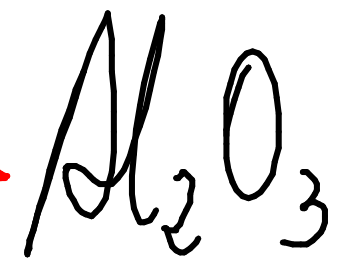
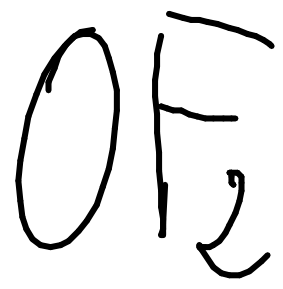
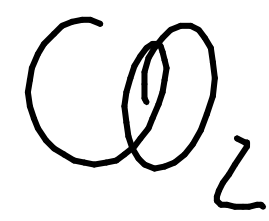
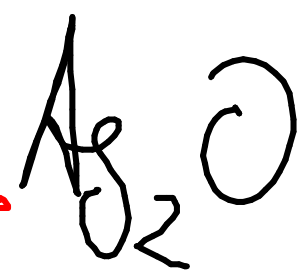
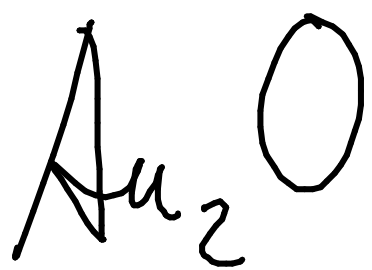
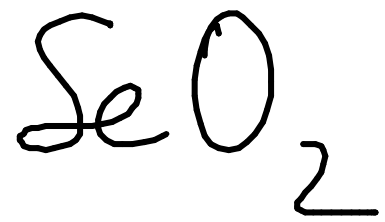
a) HF b) SbH_3 c) H_2Se d) CH_4 e) HI f) SiH_4 g) H_2S h) HCl

- a) ac. fluorhídrico ou fluoruro de hidróxeno
- b) trihidruo de antimonio ou estibamina
- c) ac. selenhídrico ou seleniuro de hidróxeno
- d) metano ou tetrahidruo de carbono
- e) ac. iodhídrico ou ioduro de hidróxeno
- f) silano ou tetrahidruo de silicio
- g) ac. sulfhídrico ou sulfuro de hidróxeno
- h) ac. clorhídrico ou cloruro de hidróxeno

7 Formula:

- a) Óxido de estaño(II)
- b) Óxido de dimercurio
- c) Pentaóxido de difósforo
- d) Óxido de plomo(IV)
- e) Óxido de plomo(II)
- f) Dióxido de manganeso





g) Dióxido de selenio

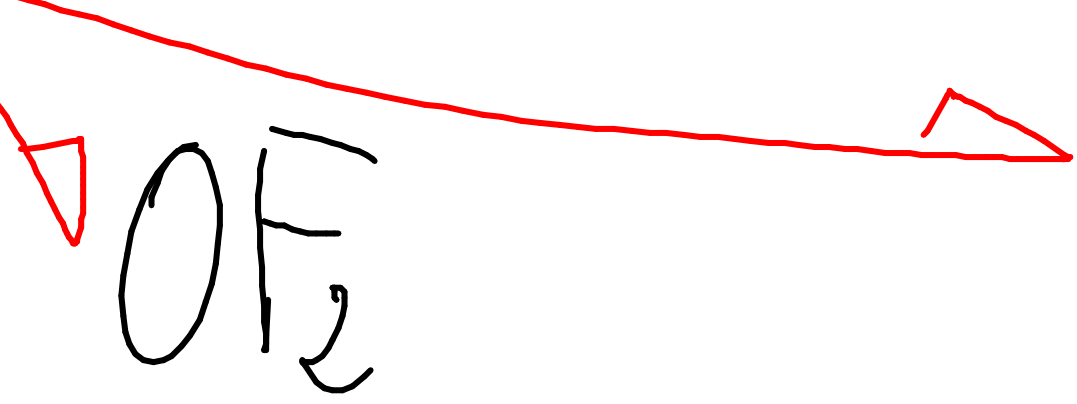
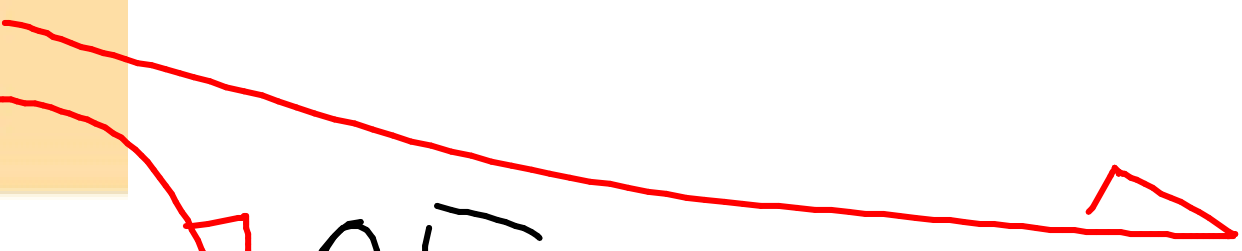
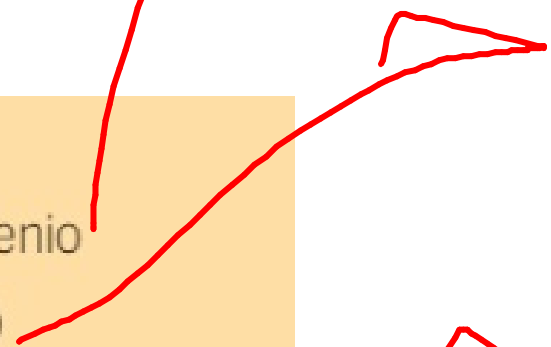
h) Óxido de oro(I)

i) Óxido de plata

j) Dióxido de carbono

k) Trióxido de dialuminio

l) Difluoruro de oxígeno



8 Nombra

- a) MnO
- b) Sb_2O_3
- c) NiO
- d) CdO
- e) SeO_3
- f) Co_2O_3

- a) óxido de manganeso (II) ou monóxido de manganeso
- b) óxido de antimónio (III) ou trióxido de diantimónio
- c) monóxido de níquel ou óxido de níquel (II)
- d) óxido de cádmio (II) ou monóxido de cádmio
- e) trióxido de selénio ou óxido de selénio (VI)
- f) trióxido de dicobalto ou óxido de cobalto (III)

g) PbO_2

h) N_2O_5

i) P_2O_3

j) CuO

k) O_3Cl_2

l) OF_2

g) dióxido de plomo ou óxido de plomo (IV)

h) pentaóxido de dinitróxeno ou óxido de nitroxeno (V)

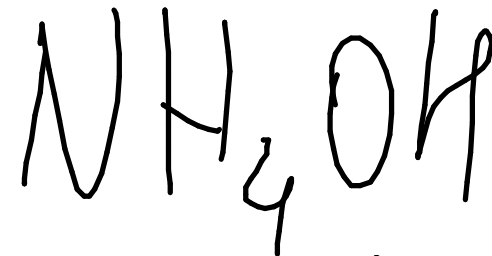
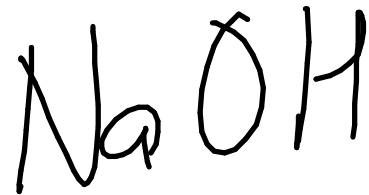
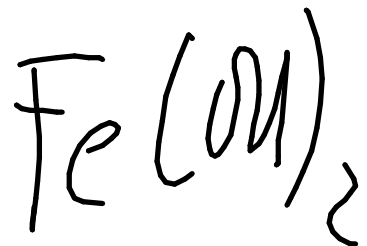
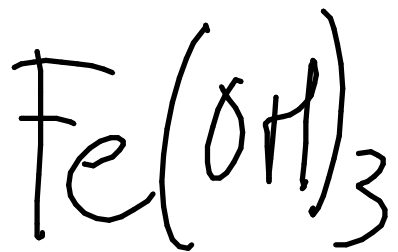
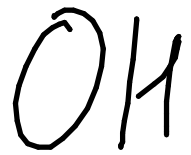
i) trióxido de difósforo ou óxido de fosforo (III)

j) monóxido de cobre ou óxido de cobre (II)

k) trióxido de dicloro ou óxido de cloro (III)

l) difluoruro de osíxeno ou fluoruro de osíxeno (II)

Hidróxidos $\begin{cases} \text{óxido O} \\ \text{hidruro H} \end{cases}$



Hidróxido
amoníaco

cal
apagada